Copy and complete the following into your notebook with the heading: Lab: <u>Heat of Reaction</u>

Question: $Mg(s) + 2HCl(aq) \text{ -----> } H_2(g) + MgCl_2(aq)$, $\Delta H = ?$

Data:

Mass of Mg _____

Mass of 1 M HCl(aq) in calorimeter _____ (≈ water)

Temperature change:

Water:

Ti = _____

Tf = _____

Energy Transfer:

<u>Water</u>

q_{H_2O} =

<u>Energy released by the reaction</u>

q_{rxn} =

<u>Heat of Reaction (enthalpy)</u>

$\Delta H = q_{rxn}$ /mol Mg =

Final Answer:

$Mg(s) + 2HCl(aq) \text{ -----> } H_2(g) + MgCl_2(aq)$, ΔH = _____

Now correctly include the heat of reaction within the chemical equation itself, as a reactant or product.

$Mg(s) + 2HCl(aq) \text{ -----> } H_2(g) + MgCl_2(aq)$

[Get Completed Stamp When Done!]