

8.1 Theories of acids and bases

Understandings:

- A Brønsted–Lowry acid is a proton/ H^+ donor and a Brønsted–Lowry base is a proton/ H^+ acceptor.
- Amphiprotic species can act as both Brønsted–Lowry acids and bases.
- A pair of species differing by a single proton is called a conjugate acid-base pair.

Applications and skills:

- Deduction of the Brønsted–Lowry acid and base in a chemical reaction.
- Deduction of the conjugate acid or conjugate base in a chemical reaction.