Name_____ Key

WP Unit 3 Practice: Compound Naming

Part A: Name the following binary compounds. If there is a transition element, use the stock naming system (i.e., using Roman numerals).

NaCl Sodium chloride

K_2S Potassium sulfide

CaO Calcium oxide

$AlBr_3$ Aluminum bromide

FeI_2 Iron(II) iodide

Fe_2O_3 Iron(III) oxide

PCl_3 phosphorus trichloride

PCl_5 phosphorus pentachloride

NO_2 Nitrogen dioxide

CuS copper(II) sulfide

Cu_2S copper(I) sulfide

S_2Cl_2 Disulfur dichloride

$TiCl_4$ Titanium(IV) chloride

Part B: Write the formulas for the following binary compounds.

calcium oxide CaO

sodium selenide Na_2Se

Magnesium nitride Mg_3N_2

Lithium bromide LiBr

(over)

Cesium chloride CsCl

Sulfur hexafluoride SF_6

Barium sulfide BaS

Mercury(II) sulfide HgS

Dinitrogen tetraoxide N_2O_4

Iron(III) oxide Fe_2O_3

Tin(II) iodide SnI_2

Tin(IV) oxide SnO_2

Write the ~~formulas~~ names for each of the following acids:

$HCl_{(aq)}$ Hydrochloric acid

$HNO_{3(aq)}$ Nitric acid

$HClO_4$ (aq) Perchloric acid

H_2S (aq) Hydrosulfuric acid

$HClO_3$ (aq) chloric acid

Write the ~~names~~ formulas for each of the following acids:

sulfuric acid H_2SO_4

carbonic acid H_2CO_3

Hydrobromic acid HBr

Nitric acid HNO_3

Acetic acid $HC_2H_3O_2$ (CH_3COOH)