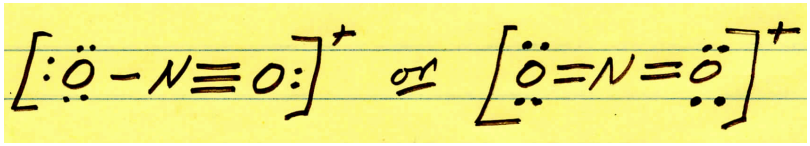
Complete the following assignment in your class notebook with the heading: Covalent compounds

1.) Draw the resonance structures for each of the following compounds:

a. O_3 c. SeO_2

b. CO_3^{2-} d. CHO_2^-

2.) Use the concept of formal charge to determine the most likely structure for NO_2^+,



3.) Use the table of bond energies below to calculate the energy needed to break all the bonds in the following:

a. CH_3OH b. H_2CO_3 (an oxyacid) c. CH_3NH_2

d. C_2BrH e. CHO_2^-

Bond Energies

Bond	Average bond energy (kJ/mol)	Bond	Average bond energy (kJ/mol)	Bond	Average bond energy (kJ/mol)	Bond	Average bond energy (kJ/mo
H–H	436	C–C	346	C–C	346	C–O	358
F–F	159	C–N	305	C=C	612	C=O	732
Cl–Cl	243	C–O	358	C≡C	835	C≡O	1072
Br–Br	193	C–H	418	C–N	305	N–N	163
I–I	151	C–Cl	327	C=N	615	N=N	418
H–F	569	C–Br	285	C≡N	887	N≡N	945
H–Cl	432	N–N	163				
H–Br	366	N–H	386				
H–I	299	O–H	459				